

# Metals and their Compounds

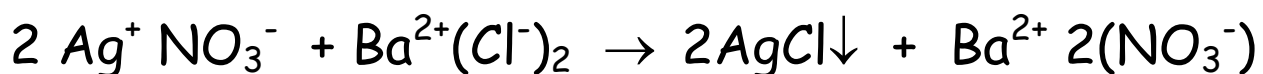
## Lecture 1.5

(ION sizes - dont need to remember this ! )

	Li <sup>+</sup>	Na <sup>+</sup>	K <sup>+</sup>	Rb <sup>+</sup>	Cs <sup>+</sup>
Angstroms 10 <sup>-8</sup> cm	0.6	0.95	1.33	1.48	1.69
	Be <sup>2+</sup>	Mg <sup>2+</sup>	Ca <sup>2+</sup>	Sr <sup>2+</sup>	Ba <sup>2+</sup>
	0.31	0.65	0.99	1.13	1.69

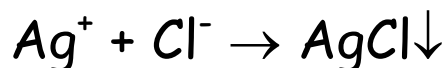
Chemistry of active metals is dull.

Ions are usually *spectator ions* - example Lab experiment silver nitrate plus barium chloride



The Ba<sup>2+</sup> (and NO<sub>3</sub><sup>-</sup>) ions play *no part* in the reaction - needed as *counterions*.

When the above equation is written as



it hides this fact.